







Is aniline soluble in hcl. The hydrolysis constant of aniline hydrochloride by a hydrogen ion method. Aniline hydrochloride formula. Why does aniline dissolve in hcl. Determination of hydrolysis constant of aniline hydrochloride. Aniline hydrochloride structure.

Calculate the hydrolysis constant and the degree of hydrolysis of aniline hydrochloride ! Posted on 2013-12-17 16:12:00 in Chemistry by ANTONIO In an experiment to determine the hydrolysis constant of aniline hydrochloride at 200C, 15 g of the salt were shaken with 1000 cm<sup>3</sup>of water and 100 cm<sup>3</sup> of benzene until equilibrium was established. 50 cm<sup>3</sup> of benzene layer was then found to contain 0.2360 g of aniline. The partition coefficient of aniline between benzene and water at 200C is found to be 11.3. Calculate the hydrolysis constant and the degree of hydrolysis of aniline hydrochloride. No Answers Posted Yet. Tardigrade - CET NEET JEE Exam App © 2022 Tardigrade®. All rights reserved Text SolutionAnswer :  $h = 2.12 \times 10^{-2}$ ,  $K(h) = 1.43 \times 10^{-5}$ M Solution :  $E_{\text{cell}} = -0.0591 \log \frac{[\text{H}^+][\text{Anode}]}{[\text{H}^+][\text{Cathode}]}$   
 $-0.188 = -0.0591 \log \frac{1}{[\text{H}^+]} \dots$   
 $\text{pH} = 3.18 = 7 - \frac{1}{2} \text{p}^{\circ}(\text{K}_b) - \frac{1}{2} \log C$   
 $3.18 = 7 - \frac{1}{2} \text{p}^{\circ}(\text{K}_b) - \frac{1}{2} \log(1/32)$   
 $\text{K}_b = 7.15 \times 10^{-10}$   
 $\text{K}_b = \frac{k(w)}{k(b)} = \frac{(10^{-14})}{(7.15 \times 10^{-10})} = 1.39 \times 10^{-5}$   
 $h = \sqrt{\frac{\text{K}_b}{\text{C}}} = \sqrt{\frac{1.39 \times 10^{-5}}{1/31}} = 2.1 \times 10^{-2}$  DISCUSSION This experiment was conducted based on three purposes. Firstly, to introduce some important types of electrodes in electrochemical and environmental studies. Secondly, to show how to use glass electrodes in the determination of ionization constants of weak poly acid through titration of pH. Then, to study the reaction of hydrolysis using glass electrodes. For this experiment, there were two parts which were determination of ionization constants of phosphoric acid in part A and determination of degree and constant of hydrolysis of aniline hydrochloride in part B. To make the electrochemical cell, there were a few types of electrode which used. There were, gas electrode, insoluble salt-metal electrode, metal electrode and membrane electrode. Basically,...show more content...The aniline hydrochloride was given in the solid form. So, the mass of aniline hydrochloride to prepare 0.1 M of aniline hydrochloride need to be calculated. The only information given was the molar mass of the aniline hydrochloride. The mole of the of the aniline hydrochloride was calculated by using formula, no. of mole = (molarity x volume)/1000. The mole of aniline hydrochloride calculated is 0.01 mol. Then, by using o. of mole = mass/(molar mass), the mass obtained is 1.296 g. 1.296 g of aniline hydrochloride was put into the 100 ml volumetric flask and distilled water was filled to the calibration mark. To determine the volume of aniline hydrochloride needed to be dissolved to prepare the solution as instructed, the formula of M1V1 = M2V2. M1 and V1 are the molarity and volume of the concentrated stock solution while M2 and V2 are the molarity and volume of the diluted solution which need to be prepared. For each of the concentration which need to be prepared, the calculated volume of aniline hydrochloride was put into the volumetric flask and distilled water was filled to the calibration mark. Verified Hint: We can use two formulas to solve this problem as: (i)  $E_{\text{cell}} = -0.0591 \log \frac{[\text{H}^+][\text{Anode}]}{[\text{H}^+][\text{Cathode}]}$  which will be used to find the emf of the cell and  $\text{pH} = 7 - \frac{1}{2} \text{p}^{\circ}(\text{K}_b) - \frac{1}{2} \log C$  which will be used to find the pH. Complete step-by-step answer: We know that to calculate the emf of the cell we can use the formula:  $E_{\text{cell}} = E_{\text{cathode}} - \frac{0.0591}{n} \log \frac{[\text{H}^+][\text{Anode}]}{[\text{H}^+][\text{Cathode}]}$  Where  $E_{\text{cell}}$  is the standard potential of the cell. In the question given above, the value of standard emf will be zero. So, the formula will be:  $E_{\text{cell}} = -\frac{0.0591}{n} \log \frac{[\text{H}^+][\text{Anode}]}{[\text{H}^+][\text{Cathode}]}$  Given the concentration of the hydrogen ions in the anode side is 1 and the value of the emf of the cell is -0.188 V, putting these values in the formula, we get:  $-0.188 = -0.0591 \log \frac{1}{[\text{H}^+]}$  On solving this, we get:  $\log \frac{1}{[\text{H}^+]} = 3.18$  We know that the value of pH of the solution is equal to the negative logarithm of the concentration of the hydrogen ions. So, we can call this value as the pH,  $\text{pH} = 3.18$  For aniline hydrochloride, the formula for pH of the solution will be:  $\text{pH} = 7 - \frac{1}{2} \text{p}^{\circ}(\text{K}_b) - \frac{1}{2} \log C$  The given concentration is  $\frac{1}{32}$  and pH is 3.18, we can calculate the  $\text{p}^{\circ}(\text{K}_b)$ . We can write:  $3.18 = 7 - \frac{1}{2} \text{p}^{\circ}(\text{K}_b) - \frac{1}{2} \log \left( \frac{1}{32} \right)$   $\text{p}^{\circ}(\text{K}_b) = 7.15$   $\text{p}^{\circ}(\text{K}_b) = 7.15$   $\text{p}^{\circ}(\text{K}_b) = -\log \text{K}_b$   $\text{K}_b = 10^{-7.15} = 7.15 \times 10^{-8}$  Now, we can calculate the value of  $\text{K}_b$  by taking the formula:  $\text{K}_b = \frac{[\text{K}^+][\text{OH}^-]}{[\text{C}]} = \frac{[\text{K}^+][\text{OH}^-]}{[\text{C}]}$  The value of  $\text{K}_b$  is  $7.15 \times 10^{-8}$  Now putting the value in the formula, we get:  $\text{K}_b = \frac{[\text{K}^+][\text{OH}^-]}{[\text{C}]}$   $7.15 \times 10^{-8} = \frac{[\text{K}^+][\text{OH}^-]}{[\text{C}]}$   $[\text{K}^+][\text{OH}^-] = 7.15 \times 10^{-8} [\text{C}]$   $[\text{K}^+][\text{OH}^-] = 7.15 \times 10^{-8} \times \frac{1}{32}$   $[\text{K}^+][\text{OH}^-] = 2.1 \times 10^{-9}$  So, the hydrolysis constant is  $2.1 \times 10^{-9}$  Therefore, the correct answer is option (b). Note: We have used to value of pH of the aniline hydrochloride as  $\text{pH} = 7 - \frac{1}{2} \text{p}^{\circ}(\text{K}_b) - \frac{1}{2} \log C$  because the aniline hydrochloride is a weak acid. For weak bases we can use  $\text{pH} = 7 + \frac{1}{2} \text{p}^{\circ}(\text{K}_a) + \frac{1}{2} \log C$ .

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